



FILE COPY

**CERTIFICATE OF INCORPORATION
ON CHANGE OF NAME**

Company No. 6730544

The Registrar of Companies for England and Wales hereby certifies that

VESTBROOK INVESTMENTS LIMITED

having changed its name, is now incorporated under the name of

TEENYTOTS LIMITED

Given at Companies House on **14th July 2009**



C06730544D



Companies House
— for the record —



THE OFFICIAL SEAL OF THE
REGISTRAR OF COMPANIES

100002/10

COMPANIES ACTS
WRITTEN RESOLUTION
OF

NAME OF EXISTING COMPANY: VESTBROOK INVESTMENTS LIMITED

COMPANY NUMBER: 6730544

We the undersigned, being all the members of the above company, for the time being entitled to receive notice of, attend and vote at General Meetings, hereby unanimously pass the following resolution and agree that the said resolution shall for all purposes be as valid and effective as if the same had been passed at a General Meeting of the company duly convened and held.

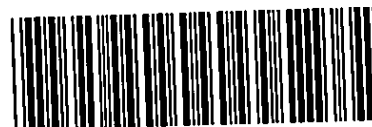
It is resolved that the name of the company be changed to:~

TEENYTOTS LIMITED

Dated this 9th day of July 2009.

Signed M. Prasad (MICHAEL PRASAD)

SATURDAY



A46 11/07/2009 24
COMPANIES HOUSE

Notes :

- The resolution must be delivered to Companies House within 15 days of it being passed.
- A £10 fee is required to change the name (cheques made payable to "Companies House").
- Have you checked whether the new name is available at www.companieshouse.gov.uk ?
- Please provide the name and address to which the certificate is to be sent.

THEORY

1. THEORY OF THE EXPERIMENT

1.1

The first part of the experiment is the determination of the equilibrium constant for the reaction:



The equilibrium constant for this reaction is defined as the ratio of the product of the concentrations of the products to the product of the concentrations of the reactants, each raised to the power of its stoichiometric coefficient. For the reaction (1), the equilibrium constant is given by:

$$K_{\text{eq}} = \frac{[\text{FeSCN}^{2+}]}{[\text{Fe}^{3+}][\text{SCN}^{-}]} \quad (2)$$

The second part of the experiment is the determination of the equilibrium constant for the reaction:



The equilibrium constant for this reaction is defined as the ratio of the product of the concentrations of the products to the product of the concentrations of the reactants, each raised to the power of its stoichiometric coefficient. For the reaction (3), the equilibrium constant is given by:

$$K_{\text{eq}} = \frac{[\text{Fe(OH)}_3]}{[\text{Fe}^{3+}][\text{OH}^{-}]^3} \quad (4)$$

The third part of the experiment is the determination of the equilibrium constant for the reaction:

The fourth part of the experiment is the determination of the equilibrium constant for the reaction:

$$K_{\text{eq}} = \frac{[\text{FeSCN}^{2+}]}{[\text{Fe}^{3+}][\text{SCN}^{-}]} \quad (5)$$