

Company Number 5573528

**1-53 Harewood Terrace Freehold (Southall) Limited**  
**Accounts for the year to 30<sup>th</sup> September 2009**  
**Balance Sheet**

Balance Sheet as at	30.9.2009	30.9.2008
	£	£
<b>Fixed Asset at Cost</b>	59,000	59,000
<b>Assets</b>		
Agents Current Account	9,095	nil
Sundry Debtors	22	2,570
<b>Liabilities</b>		
Creditors falling due within one year	nil	nil
Creditors falling due after one year	83,500	76,951
	<u>(£15,383)</u>	<u>(£15,381)</u>
<b>represented by</b>		
Authorised Share Capital 20 shares at £1	<u>20</u>	<u>20</u>
Issued Share Capital	20	20
Profit and Loss Account	(15,363)	(15,401)
	<u>(£15,383)</u>	<u>(£15,381)</u>

- (a) For the year ending 30<sup>th</sup> September 2009 the company was entitled to exemption from audit under section 477(2) of the Companies Act 2006.
- (b) The members have not required the company to obtain an audit in accordance with section 476 of the Companies Act 2006.
- (c) The directors acknowledge their responsibility for complying with the requirements of The Companies Act 2006 with respect to accounting records and the preparation of accounts.
- (d) The accounts have been prepared in accordance with the special provision in part VII of the Companies Act 1985 relating to small companies.

Approved by the board of Directors on \_\_\_\_ October 2009.

and signed on their behalf by Margaret Proctor (Director) *Margaret Proctor* 11.11.09

THURSDAY



PC2 19/11/2009 592  
COMPANIES HOUSE

THE UNIVERSITY OF CHICAGO  
DEPARTMENT OF CHEMISTRY  
RESEARCH REPORT

REPORT NO. 1000

DATE: 1960

BY: J. H. DUNN

TO: THE DEPARTMENT OF CHEMISTRY

CHICAGO, ILL.

1. TITLE: THE EFFECT OF TEMPERATURE ON THE RATE OF REACTION OF

HYDROGEN PEROXIDE WITH FERROUS SULFATE

2. SUMMARY: The rate of reaction of hydrogen peroxide with ferrous sulfate was studied at various temperatures. The results show that the rate of reaction increases with increasing temperature.

3. INTRODUCTION: The reaction of hydrogen peroxide with ferrous sulfate is a well-known reaction. It is a redox reaction in which the ferrous ion is oxidized to the ferric ion and the hydrogen peroxide is reduced to water.

4. EXPERIMENTAL: The reaction was studied by measuring the rate of disappearance of ferrous sulfate. The reaction was carried out in a series of flasks at different temperatures. The concentration of hydrogen peroxide was kept constant.

5. RESULTS: The results of the experiment are shown in the following table:

6. DISCUSSION: The results of the experiment show that the rate of reaction increases with increasing temperature. This is in agreement with the general principle that the rate of reaction increases with increasing temperature.

7. CONCLUSION: The rate of reaction of hydrogen peroxide with ferrous sulfate increases with increasing temperature.

8. REFERENCES: 1. J. H. Dunn, *Journal of Chemical Education*, 37, 100 (1960).

9. ACKNOWLEDGMENT: This work was supported by the National Science Foundation.

10. DISTRIBUTION: This report is available from the University of Chicago Press.

11. FOOTNOTES: 1. The reaction was studied at 25°C, 30°C, 35°C, 40°C, and 45°C.

2. The concentration of hydrogen peroxide was 0.01 M.

3. The concentration of ferrous sulfate was 0.01 M.

4. The reaction was studied in a series of flasks.

5. The rate of reaction was measured by the disappearance of ferrous sulfate.

6. The results of the experiment are shown in the following table:

7. The rate of reaction increases with increasing temperature.

8. This is in agreement with the general principle that the rate of reaction increases with increasing temperature.

9. The rate of reaction of hydrogen peroxide with ferrous sulfate increases with increasing temperature.

10. This work was supported by the National Science Foundation.

11. This report is available from the University of Chicago Press.